

CONJUGATED LINOLEIC ACID ALKYL ESTERS  
IN FEEDSTUFFS AND FOOD

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**FIELD OF THE INVENTION**

This invention relates to the field of nutrition  
10 and the supplementation of feedstuffs and food with alkyl esters of conjugated linoleic acid. A process for making the esters utilizes a nonaqueous alcoholate catalyzed reaction to convert the alkyl esters of linoleic acid derived from sunflower and safflower oil  
15 to predominately the c9,t11- and t10,c12- conjugated isomers.

**BACKGROUND OF THE INVENTION**

The biological activity of conjugated linoleic acids (hereinafter CLA) has been well documented in a number of indications. Its effect as an  
20 anticarcinogenic agent was demonstrated in a rat mammary tumor model by Ha, et al., Cancer Res., 52: 2035s (1992), and in a mouse forestomach neoplasia model (Ha, et al., Cancer Res., 50: 1097 (1990)). CLA has been found effective in attenuating allergic  
25 reactions mediated by type I or TgE hypersensitivity. As a nutritive supplement, CLA administration results in selective reduction in body fat, as disclosed in  
30 U.S. Patent No. 5,554,646, and has a significant positive effect on feed conversion efficient as shown in U.S. Patent No. 5,428,072.

Linoleic acid is an important component of  
35 biolipids, and comprises a significant proportion of triglycerides and phospholipids. It is an essential fatty acid, in that it is required in the diet for maintenance of healthy cells, but the body does not possess the enzymatic machinery to synthesis the fatty

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acid itself. Linoleic acid has 18 carbon atoms with double bonds at positions 9 and 12. The conjugated forms of linoleic acid have the double bond positions shifted so that the double bond pairs are separated by a single methylene group. The rearrangement of the double bonds of linoleic acid to conjugated positions results in eight possible geometric isomers of 9,11 and 10,12 octadecanoic acid (c9,c11; c9,t11; t9,c11; t9,t11; c10,c12; c10,t12; t10,c12; and t10,t12). Other minor conjugated forms appear in nature and result from synthetic processes, namely, ct8,ct10 and ct11,ct13 being the most prevalent.

A general mechanism for the isomerization of linoleic acid was described by J.C. Cowan in JAOCs, 72: 492 (1950). It is believed that the double bond is polarized by the result of a collision with an activating catalyst. The polarized carbon atom and its adjoining carbon are then free to rotate and the forces are such to make the deficient carbon atom essentially planar. When the system relieves forces set up as a result of the molecular collision, both cis and trans isomers are formed. More of the 10,12 and 9,11 isomers are formed than other species because of the thermodynamic stability of these forms. More severe conditions of heat, pressure, and polarity tend to drive isomerization further to the more stable trans,trans isomers, and cause redistribution of the double bonds with the appearance of significant quantities of the 8,10 and 11,13 forms.

One problem with aqueous alkali isomerization, which is the principal industrial process for producing CLA, is the formation of these multiple species. The reaction becomes uncontrolled and a significant proportion of the linoleic acid substrate is sacrificed to undesirable trans, trans isomers. For industrial use in drying oils where generalized polymerization between fatty acid strands is sought, it makes little difference which species of conjugated isomer

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predominate. However, in therapeutic or nutritional applications, the t10,c12 and c9,t11 isomers are believed to contain most, if not all, of the biological activity.

5 Other methods have been described utilizing metal catalysts, which are not highly efficient. Isomerization in these systems could be achieved more rapidly in the presence of higher molecular weight solvents. Kass, et al., J. Am. Chem. Soc., 61: 4829  
10 (1939) showed that replacement of ethanol with ethylene glycol resulted in both an increase in conjugation in less time. U.S. Patent No. 2,350,583 and British Patent No. 558,881 (1944) achieved conjugation by reacting fatty acid soaps of an oil with an excess of  
15 aqueous alkali at 200-230 degrees C. and increased pressure. Among the processes known to cause isomerization in the absence of aqueous alkali, is a nickel-carbon catalytic method, as described in Radlove, et al., Ind. Eng. Chem., 38: 997 (1946).  
20 Processes have also been described for isomerization of polyethanoid fatty acids in their ester forms. U.S. Patent Nos. 2,242,230 and 3,162,658 disclose methods in which the lower alkyl esters of linoleic acid are isomerized by catalysis with basic  
25 alcoholates, preferably sodium or potassium at moderate temperatures in the range of 100-140 degrees C. Typically these processes are used to generate industrial drying oils, and hence predominately utilize soy and corn oil as the starting material, in order to enhance polymerization when coated onto surfaces.  
30 These fatty acid ester compositions are not suitable for human or animal consumption because of high phosphidyl and other residue content. Purification by distillation, differential extraction, and the like removes the residues, but also causes further double bond rearrangements giving an unacceptable level of trans,trans CLA isomers, and intermolecular polymers.  
35 The purified CLA utilized in prior feeding studies

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was obtained by small scale laboratory procedures involving production of CLA from highly purified linoleic acid. For example, Sullivan, J. Am. Oil Chemists' Soc., 53: 359 (1976) describes a laboratory semi-pilot steam refining system made entirely of glass. While such systems are adequate for producing quantities of CLA for laboratory studies, or even clinical trials, they are not suitable for commercial scale bulk production. On the other hand, the large scale systems available to produce industrial quantities of CLA cannot be run inexpensively enough to produce material for bulk animal feeds. The degumming, refining, and dehydration steps necessary to obtain nutritionally safe edible CLA for livestock feeding are prohibitively complex and expensive.

Economical CLA-ester production in commercial quantities is a desirable objective in light of nutritional benefits observed on a laboratory scale. The advantages of an ester derivative rather than the free CLA fatty acids include resistance to oxidation, ease of manufacture according to the process of the present invention, palatability, and compatibility with lipid feed components.

SUMMARY OF THE INVENTION

In the present invention, a feed safe conjugated linoleic acid alkyl ester is manufactured under conditions preferentially controlling isomerization to the desired 10,12 and 9,11 isomers, while limiting formation of 8,10; 11,13; and trans,trans species. Such conditions are met by employing an alkali alcoholate catalyzed reaction in which a seed oil is split to release free fatty acids from a glycerol backbone and then esterifying prior to isomerization. The key to an adaptation of this process to a commercially viable product is reduction in the process steps which add cost. Typically, residues derived from

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non-oil components of seed oils, such as sterols and phosphatides, foul equipment and reduce palatability for feed or food use. In the case of typical seed oils such as soy or corn these residues are present in sufficient quantity that a CLA-ester product could not be used in consumable products.

In the composition of the present invention, non-oil residues are not purified away from the oil component, but rather the source of oil is selected to maintain such residues at acceptable levels. By selecting safflower or sunflower oil as an oil source, critical residue levels can be controlled to between 0.1 and 0.5% phosphatides, and to an unsaponifiable sterol fraction containing between 5 and less than 20 percent each of campesterol and stigmasterol, without extensive degumming and distillation processing steps. The resulting linoleic acid alkyl ester is comprised of at least 50 percent up to about 99 percent by weight of octadecanoic acid ester isomers representing combinations of various possible individual percentages of c9,t11-octadecanoic acid alkyl ester and t10,c12-octadecanoic acid alkyl ester. In the alkali alcoholate catalyzed process roughly equal amounts of each of these ester isomers are produced, but the relative percentages can be altered by addition of one or the other of a composition enriched for one isomer. The CLA ester may then be incorporated into an animal feed by compounding the feed from conventional ingredients in a ration typical for the species and age of the animal, and blending therewith the conjugated linoleic acid alkyl esters in a biologically active concentration, generally about 0.05 to 3.5 percent by weight.

The CLA-ester product of the present invention is obtained by direct isomerization of an unrefined linoleic acid, e.g. a linoleic acid source not subjected to refining steps. The CLA-ester composition has one part comprising at least 50 percent by weight

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of ester isomers (up to substantially 100 percent) of a mixture of ester isomers of c9,t11-octadecanoic acid ester and t10,c12 -octadecanoic acid ester, a second part comprising less than about 10 percent by aggregate weight of ester isomers of the structure 8,10-octadecanoic acid ester, 11,13-octadecanoic acid ester, and trans,trans-octadecanoic acid esters, and a third part containing a phosphatidyl residue of between 0.1 and 0.5 percent of the total composition weight. The alkyl groups may be methyl, ethyl, propyl, isopropyl, butyl, isobutyl and the like. Adjustments in concentration of the c9,t11 and t10,c12 isomers can be made by addition of a composition enriched for one or the other isomer to yield an ester composition wherein the c9,t11, or the t10,c12 respectively contained in the first composition part constitutes greater than 60 percent of the total isomers of octadecanoic acid esters.

In the process embodiment of the present invention resulting in a food grade composition suitable for an animal feed, food ingredient, or human dietary supplement, an unrefined CLA-ester having a phosphatidyl residue less than 0.5 percent is treated with an alkali alcoholate in the presence of a monohydric low molecular weight alcohol such as methyl (or ethyl alcohol, continuing the treatment at low temperature (about 90 to 145 degrees C.) until at least 50 percent of the ester is converted to CLA-ester, acidifying by addition of an aqueous acid, and then separating the CLA-ester from the aqueous acid without a distillation step. ) )

#### BRIEF DESCRIPTION OF THE DRAWINGS

Figure 1 is a flow diagram of the present process, in which esterification is followed directly by isomerization without a purification step.

Figure 2 is a rectilinear plot showing the relation between aqueous alkali processing temperature and the yield of the 9,11 and 10,12 isomers compared to total conjugated linoleic acids.

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#### DETAILED DESCRIPTION OF THE PREFERRED EMBODIMENTS

Various methods of producing conjugated double bonds by alkali isomerization are known in the art. 10 U.S. Patent 2,350,583 (Bradley, 1944) describes a method of producing conjugated fatty acids by aqueous alkali isomerization. This method resulted in the conjugation of about 50% of the double bonds present in the polyunsaturated fatty acids used. U.S. Patent 15 2,242,230 (Burr et al., 1941) describes a method of non-aqueous alkali conjugation of fatty acids, resulting in the conjugation of approximately 100 percent of the double bonds in the polyunsaturated fatty acids studied. Another process resulting in the 20 efficient formation of conjugated double bonds is described in U.S. Patent 4,381,264 (Struve, 1983). There, the inventors treat polyunsaturated fatty acids with  $\text{SO}_2$  in the presence of substoichiometric amounts 25 of soap forming bases. Perhaps the most commercially viable method for producing large quantities of conjugated fatty acids is the continuous flow aqueous alkali isomerization process described in U.S. Patent 4,164,505. This process results in essentially all available double bonds being conjugated in a short 30 reaction time. The foregoing patents are incorporated herein by reference.

In the production of a food grade CLA-ester, a laboratory scale procedure utilizes reagent grade 9,12-linoleic acid alkyl ester in a reaction as disclosed in 35 U.S. Patent Nos. 2,242,230 and 3,162,658, hereby incorporated by reference. The reaction is carried out at about 100-140 degrees C. in a closed vessel for 5-8 hours. Isomerization is substantially complete. The

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resulting product is perfectly safe for human or animal consumption. However, the use of reagent grade linoleic acid esters is impractical for compounding into bulk feeds, or even for human consumption in capsule form. On the other hand mere isomerization of the linoleic acid contained in a crude or unrefined linoleic ester, as is suitable for production of the industrial drying oils contemplated by the above cited patents, will not yield a food grade because of the high phosphatidyl and sterol content, as well as a tendency for oils containing high levels of these impurities to polymerize.

Applicants have discovered that if sunflower or safflower oil is utilized, the resulting phosphatidyl and sterol residues are low enough for the material to be characterized as food grade. In particular, Applicants first provide the sunflower or safflower as delivered by tanker, subject it to fat slitting and esterification, and then proceed directly with the unrefined esterification product into isomerization. The usual purification and refining steps are not necessary. This material can be made in commercial scale quantities at a low enough cost to be practical as a feed or food supplement.

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Referring now to the flow diagram of Figure 1, after fat splitting and dehydration, the free fatty acids are combined with methanol or other monohydric low molecular weight alcohol and heated to the temperature at which the alcohol boils. Esterification proceeds under refluxing conditions with removal of the reaction water through a condenser. After the addition of a further quantity of the same or a different monohydric alcohol an alcoholeate catalyst is blended into the ester mix. Typical alcoholeate catalysts are sodium or potassium ethoxide, or their methyl, butyl, or propyl counterparts.

In the esterification, methanol or ethanol are preferred, although other branched or straight chain

monohydric alcohols may be used. The longer the aliphatic chain of the alkyl group, the more lipid compatible the material becomes. Also the viscosity tends to increase. For different types of feed or food, whose consistency varies, product of varying viscosity can be used to obtain the desired flow or compounding characteristics without affecting the therapeutic or nutritional properties arising from the CLA moieties. The theory and practice of esterification are conventional. A basic explanation of the most common methods is set forth in the McCraw-Hill Encyclopedia of Science & Technology, McGraw-Hill Book Co., N.Y.: 1996 (5th ed.). The animal and human body has a variety of esterases, so that the CLA-ester is cleaved to release the free fatty acids readily. Tissue uptake may have a different kinetics depending on the tissue involved and the benefit sought.

In the isomerization step, it was found that alcoholate catalysis produced a much superior product than aqueous alkali mediated isomerization. The latter process always produced undesirable isomers even under mild reaction conditions. The milder conditions do give lower amounts of unwanted isomers, but at the great expense of yield, as shown in the Examples. In most systems the appearance of the c9,t11 and t10,c12 isomers dominates and they are formed in roughly equimolar amounts. It has not heretofore been possible to control the isomerization of the one isomer to the exclusion of the other. While it is desirable to increase the percentage of one or the other isomer (depending on the physiological effect to be achieved), at present this must largely be carried out by adding an enriched source of the desired isomer.

Crude sunflower or safflower oil is the preferred fatty acid source for producing CLA. Sunflower oil contains a high amount of linoleic acid (about 65% on average). Safflower oil typically contains even higher amounts (greater than 70%). Preferably, a hexane

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extract of crude, non-degummed oil is the starting substrate for CFAP production. This extract is commercially available and is the same quality as the oil used as the starting point for edible products.

5 The ability to use raw sunflower or safflower oil as the starting substrate provides an important economic advantage because it is less expensive than refined sunflower oil.

Applicants have discovered that other raw oils, such as raw corn and soybean oils, are not suited to the present new use of CLA in bulk feeds because of the production of polymerized products during the fat splitting and conjugation processes, and because of the high phosphatide content. Also, certain sterols such as campesterol and stigmasterol are known to have a tendency to foul processing equipment during conjugation and plug nozzles during materials transfer. The polymerization by-products also result in loss of yield from these other oils, even though at first glance the other oils may seem to have more desirable properties. Corn oil (about 56% linoleic acid) and soybean oil (about 50-55% linoleic acid) have comparable linoleic acid contents as compared to sunflower oil (about 60% linoleic acid). These oils are inexpensive and large quantities are available, which make them attractive candidates as a potential source of CLA for bulk feeds. However, their use for commercial CLA production is substantially lower per unit quantity of oil because the ultimate yields of CLA are lower than for sunflower or safflower oil, and because of the added expense for additional cleaning and purification steps.

Heat sensitive triglycerides containing multiple double bonds are abundant in oils having an iodine value above 120. Oils containing such heat-sensitive triglycerides have a tendency to form polymers when subjected to continuous countercurrent fat-splitting. These polymers become insoluble in oil and will foul

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equipment, resulting in lowered efficiencies of splitting and yields. Sunflower oils are classified as heat-sensitive because of their high linoleic acid content and iodine number. Sunflower oil has an iodine number of 130; safflower oil of about 145. Soybean oil has an iodine number of about 132, and corn oil has an iodine number of about 130. All these oils have high linoleic acid contents and iodine numbers above 120, thus belonging to the heat-sensitive group. It is therefore surprising that sunflower and especially safflower oils can be split and conjugated by commercial processes with few processing complications.

The prevalent phosphatides are phosphatidylethanolamine, phosphatidylinositol and phosphatidylcholine. The phosphatide content of sunflower and safflower oils is about 0.4 to 1.0% as reported in Klein and Crauer, JAOCS 51:382A-385A and Burkhardt, JAOCS 48:697-699 (1971), respectively. In contrast, the phosphatide content of soybean oil is about 1.5-2.5% as reported in the Handbook of Soy Oil Processing and Utilization, Erickson et al. eds., AOCS, Champaign, 1980. For crude oils with low amounts of these substances, such as sunflower and safflower oils, proceeding can advance directly to esterification and isomerization after degumming.

Example 1: Effect of Varying Temperature and Reaction Duration on CLA Yield and Composition

The effect of temperature and reaction duration on the conjugation of safflower oil was determined. Water and NaOH were added to a high pressure reactor (Parr Model 450 ML Benchtop Alloy 400, equipped with a pressure gauge and stirrer) as indicated in Table 1, columns 1 and 2. The NaOH was allowed to dissolve and safflower oil (column 3) was added to the reactor. The reactor was closed and flushed for 2 min. with nitrogen

and then all valves were closed. The reactor was heated in an electrical gasket to the desired temperature (column 4) and maintained at that temperature for the desired time (column 5). The  
5 temperature was then reduced to 60°C before pressure was released and the reactor opened. For each reaction, two grams of the resulting solidified soap were taken from the reactor and dissolved in water at approximately 40°C. Citric acid was then added to  
10 reduce the pH of the solution to below 6. A sample was withdrawn from the fatty acid top layer and prepared for Gas Chromatography.

The results of the gas chromatography are presented in column 6 (total percentage of 9,11 and  
15 10,12 isomers), column 7 (total percentage of 11,13 isomers), and column 8 (total percentage of all CLA isomers or yield). These data indicate that as reaction duration and temperature increase, the total amount of conjugation and the percentage of 11,13  
20 isomers increase. Under conditions where formation of the 11,13 isomer is low, the total amount of conjugation is also low.

Figure 2 further illustrates the above findings. As temperature is increased, total CLA also increases,  
25 but the proportion of the 9,11 and 10,12 isomers declines. At temperatures above 189°C, the absolute amounts of these isomers also declines. The CLA species favored under these conditions are the trans,trans isomers.

TABLE 1: Conjugation in Water Solvent  
Summary of Results

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	Water gram	NaOH gram	Safflower Oil gram	Mean t. °C of reaction	Time hours	9,11+ 10,12 area%	11,13 area%	CLA total area%
10	50.21	29.93	99.94	189	6.36	45.99	5.73	55.86
	70.20	29.93	99.84	187	6.40	44.94	3.23	51.28
	50.10	30.17	100.74	183	6.40	40.23	3.37	48.07
	49.91	29.93	100.40	179	6.52	32.00	1.48	34.92
15	49.97	29.80	100.02	179	10.08	41.86	3.12	48.21
	49.94	39.84	99.84	179	6.30	32.6	3.04	37.12
	29.50	24.83	99.21	240	3.25	28.37	10.78	71.58
	30.33	25.15	100.43	221	2.30	40.87	14.72	72.61
	49.92	30.00	100.36	150	6.34	7.07	0	7.44

20 Example 2: Conjugation of Safflower Fatty Acid  
Methylester (FAME)

The reaction was carried out in a closed vessel.

25 The following components were mixed together: 100g safflower FAME and a mixture of approximately 2.8 g KOCH<sub>3</sub> and 2.8 g methanol. There was probably more KOMe than methanol due to evaporation of methanol during mixing of the two components. The mixture was stirred  
30 for 5 hours at 111-115 deg C in nitrogen atmosphere in a closed reaction vessel. The distribution of isomers was analyzed by Gas Chromatography. The results are summarized in Table 2. The raw GC data is presented in Table 3. These data indicate that the conjugation  
35 safflower FAME may be accomplished under mild conditions, resulting in a product lacking appreciable amounts of undesirable 8,10 and 11,13 isomers.

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TABLE 2: Isomer Distribution

	Palmitic acid	6.6%
	Stearic acid	2.7%
5	Oleic acid	12.9%
	Linoleic acid	5.7% (unconjugated)
	CLA c9,t11	34.1%
	CLA t10,c12	33.3%
10	CLA c,c	1.8%
	CLA t,t	1.0%
	CLA total	70.2%

EXAMPLE 3: Large Scale Batch Production of Conjugated  
15 Safflower FAME

The production of safflower conjugated FAME may be divided into two steps, methanolysis and conjugation.  
20 For methanolysis, 6,000 kg safflower oil was drawn into a closed reactor. The reactor was purged with nitrogen at atmospheric pressure, and 1150 liters of methanol and 160 kg of NaOCH<sub>3</sub> (30% solution) were added. The mixture is heated to 65°C while stirring, and reacted at 65°C for 2 hours. The resulting bottom layer was  
25 decanted while the reactor was purged with nitrogen gas. 1000 liters of water (40-50°C, into which 50 kg citric acid monohydrate has been dissolved) was then added while stirring. The layers were allowed to separate (approx. 60 min.) and the bottom layer  
30 decanted while purging the reactor with nitrogen gas. The resulting safflower FAME product was dried at 80°C under vacuum for one hour.

To conjugate the safflower FAME, 250 kg of KOCH<sub>3</sub> dissolved in methanol to form a paste was added to the  
35 reactor. The mixture was then heated to 120°C while stirring and the reaction allowed to continue for 3 hours. The mixture was cooled to 100°C, and 1000 liters of water (40-50°C, into which 50 kg citric acid

monohydrate has been dissolved) was added while stirring. The mixture was stirred for 15 minutes and then the layers were allowed to separate for 20 minutes. The bottom layer was decanted and the product dried at 80°C for 1 hour and then stored under nitrogen.

5 The resulting CLA was analyzed using a Perkin Elmer Autosystem XL GC under the following conditions:

10 Column: WCOT Fused Silica 100 m X 0.25 mm,  
Coating CP-SIL 88

15 Carrier: He gas, 30.0 PSI  
Temp: 220 C  
Run time: 35-90 min.  
Inject.: Splitless, 240 C  
Detect.: FID, 280 C

The GC results are summarized in Table 3.

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TABLE 3: Gas-Chromatography Results

Peak #	Time (min)	Component Name	Area (%)	Area ( $\mu$ Vs)	Height ( $\mu$ V)
1	46.874	C16:0	6.37	29874.50	4026.29
8	58.685	C18:0	2.61	12231.70	1542.34
3	62.141	C18:1 c9	13.14	61668.78	7369.08
6	62.652		0.70	3263.62	391.92
5	66.404		0.35	1627.60	177.41
6	66.917		0.26	1239.15	157.35
7	67.583	C18:2 c9, c12	5.75	26964.95	3153.80
8	70.631		0.25	1171.90	141.41
9	75.011	CLA c9, t11	34.42	161529.90	17544.79
10	75.936	CLA t10, c12	33.48	157129.82	17157.21
11	76.400	CLA c9, c11	0.84	3935.70	302.61
12	76.631	CLA c10, c12	0.49	2316.98	279.31
13	77.905	CLA t, t 9, 11+ 10, 12	1.35	6344.50	710.88
			100.00	469299.10	52954.41

TABLE 3 cont.

5	Peak #	Time (min)	Component Name	Area (%)	Area ( $\mu$ Vs)	Height ( $\mu$ V)
	9	75.011	CLA c9,t11	48.76	161529.90	17544.79
	10	75.936	CLA t10,c12	47.43	157129.82	17157.21
	11	76.400	CLA c9,c11	1.19	3935.70	302.61
	12	76.631	CLA c10,c12	0.70	2316.98	279.31
10	13	77.905	CLA t,t 9,11+ 10,12	1.92	6344.50	710.88
				100.00	331256.90	35994.80

15       The following are examples of typical animal rations containing the CLA esters of the present invention.

EXAMPLE 4

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PIG STARTER RATIONS

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Ingredients	lbs.	kgs.
Corn, yellow (8.4% protein)	1067	484.7
Soy bean meal, solvent extracted, dehulled (47% protein)	570	259
CLA-ester	5	2.3
Whey, dried (12.0% protein)	300	136
Dicalcium phosphate	24	11
Limestone	16	7
Iodized salt	5	2
Trace mineral premix	5	2
Vitamin premix	8	4
Totals	2000	908

B

GROWER-FINISHER RATIONS FOR PIGS  
(FROM 40-240 LBS [18-109KGS])

Ingredients	lbs.	kgs.
Corn, yellow (8.4% protein)	1566	
Soybean meal, solvent extracted (44% protein)	380	
CLA-ester	5	
Dicalcium phosphate	21	
Limestone	15	
Iodized Salt	5	
Trace Mineral Premix	3	
Vitamin Premix	3	
Total	2000	

C

PIG GROWER FINISHER RATIONS  
(FOR PIGS 121-240LBS [55-109KGS])

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Ingredients	lbs.	kgs.
Corn, yellow (8.4% protein)	1687	
Soybean meal, solvent extracted (44% protein)	265	
CLA-ester	5	
Dicalcium phosphate	15	
Limestone	15	
Iodized salt	5	
Trace mineral premix	2	
Vitamin premix	3	
<b>Total</b>	<b>2000</b>	

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COMPOSITION AND ANALYSIS OF  
PIG TRACE MINERAL REMIX

25

Element	Source	Amount (lbs)
Copper (Co)	Copper Sulfate	1.500
Iodine (I)	Potassium Iodide	0.010
Iron (Fe)	Ferrous Sulfate	25.000
Manganese (Mn)	Manganese Sulfate	2.500
Selenium (Se)	Sodium Selemite)	0.025
Zinc (Zn)	Zinc Sulfate	25.000
	Carrier	45.965
<b>Total</b>		<b>100.000</b>

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DO NOT USE "DRAFT" FORM

COMPOSITION OF PIG VITAMIN PREMIX

	Vitamins	Amount
<b>Essential</b>		
5	Vitamin A.....(million IU)	5.0
	Vitamin A.....(million IU)	0.6
	Vitamin E.....(thousand IU)	26.0
10	Niacin.....(g)	25.0
	d-Pantothenic acid....(g)	20.0
	Riboflavin.....(g)	0.6
	Vitamin B-12.....(mg)	25.0
<b>Optional</b>		
15	Biotin.....(g)	0.3
	Menadione.....(g)	4.0
	Carrier.....	to 10lbs
	<b>Total</b>	<b>10.0</b>

F

18% PROTEIN LAYER RATIONS FOR HENS

	<b>Ingredients</b>	<b>lbs.</b>	<b>kgs.</b>
5	Ground yellow corn	1242	564.5
	CLA-ester	5	2.3
10	Alfalfa meal, 17%	25	11.3
	Soybean meal, dehulled	451.6	205.3
15	Meat and bone meal (47%)	50	23.0
	DL-methionine	1.0	.5
	Dicalcium phosphate	7	3.1
	Ground limestone	174	79.1
	Iodized salt	7	3.1
20	Stabilized yellow grease	37	17.2
	<b>Mineral and vitamin supplements</b>		
	Calcium pantothenate (mg)	5,000	
	Manganese (g)	52	
	Selenium (mg)	90.8	
25	Zinc (g)	16	
	Vitamin A (IU)	6,000,000	
	Vitamin D <sub>3</sub> (IU)	2,000,000	
	Choline (mg)	274,000	
	Niacin (mg)	12,000	
	Riboflavin (mg)	2,000	
	Vitamin B-12	6	
	<b>Total</b>	<b>2000</b>	<b>909.4</b>

STARTER AND FINISHER RATIONS FOR BROILERS

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	Ingredients	Starter (up to 24 days)		Finisher (25 days to market)	
		lbs.	kgs.	lbs.	kgs.
5	Ground yellow corn	1,106	503	1235	561
10	CLA-ester	5	2.3	5	2.3
15	Soybean meal, dehulled	605	275	420	191
20	Alfalfa meal, 17%	-	-	25	11
25	Corn gluten meal, 60%	50	23	75	34
30	Fish meal, herring, 65%	50	23	50	23
35	Meat and bone meal, 47%	50	23	50	23
40	Dicalcium phosphate	10	4	9	4
	Ground limestone	18	-	14	6.3
	DL-methionine	0.8	0.3	-	-
	Stabilized yellow grease	101	45.7	110	49.4
	Iodized salt	7	3	7	3
	<b>Mineral and vitamin supplement</b>				
	Calcium pantothenate (mg)	5,000		5,000	
	Manganese (g)	75		75	
	Organic arsenical supplement	0.1		0.1	
	Selenium (mg)	90.8		90.8	
	Zinc (g)	30		30	
	Vitamin A (IU)	4,000,000		4,000,000	
	Vitamin D (IU)	1,000,000		1,000,000	
	Vitamin E (mg)	2,000		2,000	
	Vitamin K (mg)	2,000		2,000	
	Choline (mg)	503,000		672,000	
	Niacin (mg)	20,000		20,000	
	Riboflavin (mg)	3,000		3,000	
	Vitamin B-12 (mg)	12		12	
	<b>Total</b>	<b>2000.9</b>	<b>909.3</b>	<b>2000.1</b>	<b>909.5</b>

GROWER/FINISHER TURKEY RATIONS

5	Ingredients	Grower (8-16 weeks)		Finisher (16 weeks-market)	
		lbs.	kgs.	lbs.	kgs.
	Ground yellow corn	1194	595	1490	677.2
	Wheat middlings	50	23	-	-
	Alfalfa meal, 17%	25	11.3	25	11.3
10	Soybean meal, dehulled	570	259	335	152.3
	Meat and bone meal, 47%	50	23	50	23
	Dicalcium phosphate	32	14.5	23	10.5
	Ground limestone	14	6	17	8
15	Stabilized yellow grease	45	20.7	45	20.7
	CLA-ester	6	2.3	5	2.3
	Iodized Salt	10	4.5	10	4.5
	<b>Mineral and vitamin supplements</b>				
20	Calcium pantothenate (mg)	4,500		4,500	
	Manganese (g)	30		30	
	Selenium (mg)	181.6		181.6	
	Zinc (g)	30		30	
25	Vitamin (IU)	1,500,000		7,500,000	
	Vitamin D (IU)	1,700,000		1,700,000	
	Vitamin E (IU)	10,000		10,000	
	Biotin (mg)	100		100	
	Choline (mg)	388,000		417,000	
30	Niacin (mg)	46,000		48,000	
	Riboflavin (mg)	5,000		5,000	
	Vitamin B-12	6		6	
	<b>Total</b>	<b>2000</b>	<b>909.3</b>	<b>2000</b>	<b>909.3</b>

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DRY DOG FOOD FORMULA

	<b>Ingredients</b>	<b>Formula 1, %</b>	<b>Formula 2, %</b>
5	Meat and bone meal, 50% CP	8.0	15.0
10	Fish meal, 60% CP, low fat	5.0	3.0
	Soybean meal, 44% CP	12.0	-
	Soybean meal, 50% CP	-	19.0
	Wheat germ meal, 25% CP	4.0	5.0
	Skimmed milk, dried	4.0	2.75
	Cereal grains, mixed	51.23	-
15	Corn, flaked	-	23.25
	Wheat bran	4.0	-
	Wheat, flaked	-	23.35
	Animal fat	1.75	2.75
	CLA-ester	.25	.25
20	Steamed bone meal	2.0	-
	Brewers yeast	2.0	5.0
	Fermentation solubles, dehydrated	1.0	-
	Salt and trace minerals	0.5	0.5
25	Vitamin mixture	0.25	0.25
	Ferric oxide	0.02	-
	<b>Total</b>	<b>100.00</b>	<b>100.00</b>

SEMI-MOIST DOG FOOD FORMULAS

	<b>Ingredients</b>	<b>Formula 1, %</b>	<b>Formula 2, %</b>
5	Soy flakes	30.9	33.5
	Meat byproducts, 70% moisture	32.0	-
	Meat and bone meal, dehydrated	-	7.3
	Water	-	25.6
10	Sugar	21.0	21.0
	Calcium and phosphorous supplement	3.3	-
	Soybean hulls	3.1	3.1
	Skimmed milk, dried	2.5	-
15	Propylene glycol	2.1	2.1
	Sorbitol	2.0	2.0
	Animal fat	.75	3.95
	CLA-ester	.25	.25
	Emulsifiers	0.9	-
20	Potassium sorbate	0.35	0.35
	Salt	0.6	0.6
	Vitamins	0.25	0.25
	<b>Total</b>	<b>100.000</b>	<b>100.000</b>